Review for Unit 7 Exam – The Stoichiometry Part

Complete the following activities and answer whatever questions are there to be better prepared for the exam Friday.

1. 35 mL of 0**.**479 M barium chloride (aq) and abundant aqueous sodium sulfate are combined. A cloudy precipitate is formed. What is that precipitate and what mass of this solid is formed?
2. Potassium bromide and lead (II) nitrate react as per the following reaction:

2KBr (aq) + Pb(NO3)2 (aq) 🡪 2KNO3(aq)+ PbBr2(s)

If 25 mL of 2M potassium bromide reacts with abundant lead (II) nitrate, what mass of potassium nitrate will be produced and remain dissolved in the solution? What mass of lead (II) bromide will precipitate out?

1. Silver nitrate and copper (II) chloride are combined and a solid precipitate cloud appears. If 35 mL of .444M CuCl2 in solution reacts with abundant silver nitrate, what mass of solid precipitates out of solution? What is that solid?