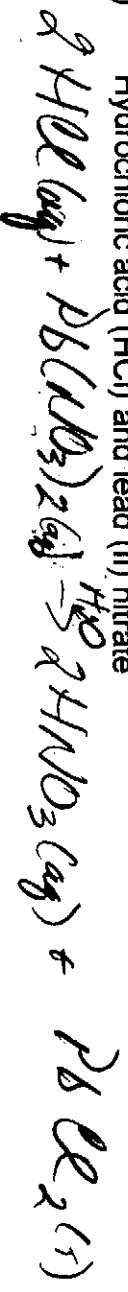


1) Using your solubility rules, write the equation for the reaction. (include aq, s, g, and l as appropriate)



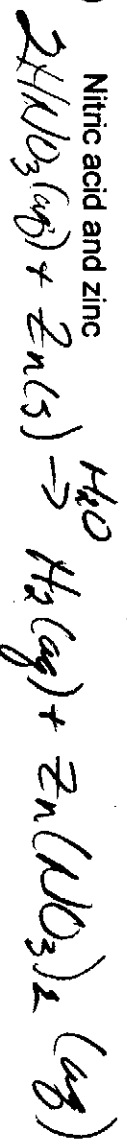
2) Hydrochloric acid (HCl) and lead (II) nitrate



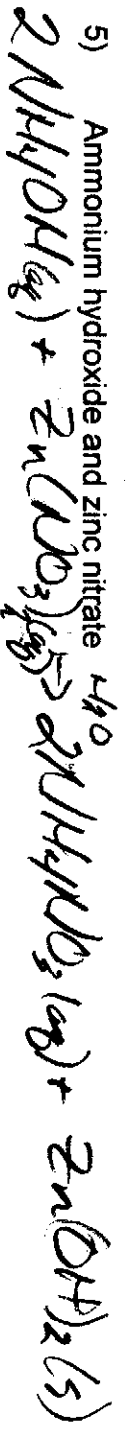
3) Sodium phosphate and calcium acetate



4) Nitric acid and zinc



5) Ammonium hydroxide and zinc nitrate

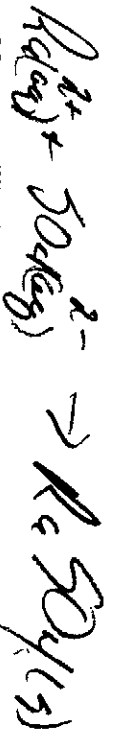


11) Write the net ionic equation for the formation of the following solids  
 example: Silver chloride is formed by  $Ag^+(aq) + Cl^-(aq) \rightarrow AgCl(s)$

6) Vanadium (IV) hydroxide



7) Radium (II) sulfate



8) Mercury (II) dichromate



9) Silver Acetate



10) Calcium sulfite

