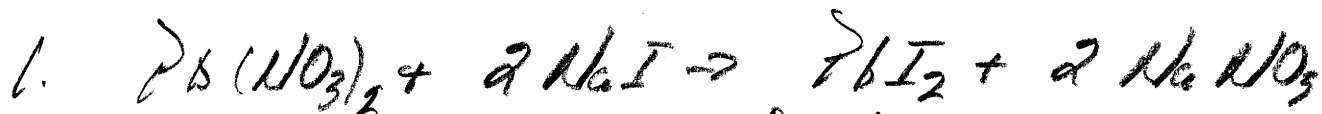


## U5 HW 1.5



$$\frac{(.206 \text{ mol NaI}) (1 \text{ mol PbI}_2)}{(2 \text{ mol NaI})} = \underline{.103 \text{ mol PbI}_2}$$

2 H-OH



$$\frac{(2.5 \text{ mol HCl}) (1 \text{ mol CaCl}_2)}{(2 \text{ mol HCl})} = \underline{1.25 \text{ mol CaCl}_2}$$



$$\textcircled{1} \frac{(7.65 \text{ g Cl}_2) (1 \text{ mol})}{(70.90 \text{ g})} = .1078 \dots \text{ mol Cl}_2$$

$$\textcircled{2} \frac{(.1078 \text{ mol Cl}_2) (2 \text{ mol NaCl})}{(1 \text{ mol Cl}_2)} = .2157 \dots \text{ mol NaCl}$$

$$\textcircled{3} \frac{(.2157 \dots \text{ mol NaCl}) (58.44 \text{ g})}{\text{mol}} = \underline{12.61 \text{ g NaCl}}$$

$$\begin{array}{r} 2.59 \\ .45 \\ \hline .44 \end{array}$$