

FITCH

10) Molecular formulas (for covalent compounds) are easy to write when given the name. Simply use the number word roots to tell you how much of each element is there. If no number word root is listed, assume there's one (mono isn't always used).

antimony tribromide $SbBr_3$ hexaboron silicide B_6Si

chlorine dioxide ClO_2 hydrogen iodide HI

iodine pentafluoride IF_5 dinitrogen trioxide N_2O_3

ammonia NH_3 phosphorus triiodide PI_3

tetraphosphorus pentasulfide P_4S_5

oxygen O_2 selenium hexafluoride SeF_6

disilicon hexafluoride Si_2F_6 sulfur tetrachloride SCl_4

methane CH_4 diboron silicide B_2Si

nitrogen trifluoride NF_3