**Formula and naming practice quiz**

Name the following compounds:

1. PO4-3 – phosphate
2. NO2 - nitrogen dioxide

3. H2SO4 – sulfuric acid

4. FeCl3 – iron (III) chloride

1. HCl – hydrochloric acid
2. Li2SO4 – lithium sulfate
3. PF5 – phosphorus pentafluoride
4. Cu(NO3)2 – copper (II) nitrate
5. KSCN – potassium thiocyanate

Write the formulas for the following compounds:

1. barium hydroxide – Ba(OH)2
2. ammonium – NH4**+**
3. Lead (IV) carbonate – Pb(CO3)2
4. Silicon tetrafluoride – SiF4
5. phosphoric acid – H3PO4
6. sodium phosphate – Na3PO4
7. aluminum sulfide – Al2S3
8. acetate – C2H3O2**-**
9. nitric acid – HNO3
10. Write the Lewis Dot Diagram for Cl and Cl-.
11. How can a person tell if they are looking at an ionic compound? A covalent (molecular) compound?

Ionic – a metal + a nonmetal / Covalent – 2 nonmetals

1. When metals ionize, they tend to form what kind of ions? Positive ion (cation)
2. Describe a polyatomic ion. An ion that is a combination of atoms.
3. How many valence electrons do the alkali metals have? 1 The alkaline earth metals? 2 The oxygen family? 6 Halogens? 7
4. Describe the octet rule. All elements try to end up with 8 valence electrons. They do this by gaining electrons, losing electrons or sharing electrons.