

Balance the following equations. Place the coefficients in the blanks to balance the equation. If the coefficient is 1 leave it blank

- $1.$ $\underline{\quad}$ $\text{N}_2 + \underline{\quad}$ $\text{O}_2 \rightarrow \underline{2}$ NO
- $2.$ $\underline{2}$ $\text{Na} + \underline{2}$ $\text{H}_2\text{O} \rightarrow \underline{2}$ $\text{NaOH} + \underline{\quad}$ H_2
- $3.$ $\underline{2}$ $\text{K} + \underline{2}$ $\text{H}_2\text{O} \rightarrow \underline{2}$ $\text{KOH} + \underline{\quad}$ H_2
- $4.$ $\underline{\quad}$ $\text{NaOH} + \underline{\quad}$ $\text{HCl} \rightarrow \underline{\quad}$ $\text{NaCl} + \underline{\quad}$ H_2O
- $5.$ $\underline{\quad}$ $\text{Fe}_2\text{O}_3 + \underline{3}$ $\text{H}_2\text{O} \rightarrow \underline{2}$ $\text{Fe}(\text{OH})_3$
- $6.$ $\underline{\quad}$ $\text{Na}_2\text{CO}_3 + \underline{2}$ $\text{H}_2\text{O} \rightarrow \underline{2}$ $\text{NaOH} + \underline{\quad}$ H_2CO_3
- $7.$ $\underline{1}$ $\text{B} + \underline{3}$ $\text{O}_2 \rightarrow \underline{2}$ B_2O_3
- $8.$ $\underline{2}$ $\text{SbCl}_3 + \underline{3}$ $\text{H}_2\text{S} \rightarrow \underline{\quad}$ $\text{Sb}_2\text{S}_3 + \underline{6}$ HCl
- $9.$ $\underline{3}$ $\text{H}_2\text{O} + \underline{\quad}$ $\text{P}_2\text{O}_5 \rightarrow \underline{2}$ H_3PO_4
- $10.$ $\underline{2}$ $\text{NaOH} + \underline{\quad}$ $\text{H}_2\text{S} \rightarrow \underline{\quad}$ $\text{Na}_2\text{S} + \underline{2}$ H_2O
- $11.$ $\underline{2}$ $\text{Fe} + \underline{3}$ $\text{H}_2\text{O} \rightarrow \underline{\quad}$ $\text{Fe}_2\text{O}_3 + \underline{3}$ H_2
- $12.$ $\underline{2}$ $\text{K}_3\text{PO}_4 + \underline{3}$ $\text{MgCl}_2 \rightarrow \underline{\quad}$ $\text{Mg}_3(\text{PO}_4)_2 + \underline{6}$ KCl
- $13.$ $\underline{2}$ $\text{CrCl}_3 + \underline{3}$ $\text{H}_2\text{SO}_4 \rightarrow \underline{\quad}$ $\text{Cr}_2(\text{SO}_4)_3 + \underline{6}$ HCl
- $14.$ $\underline{3}$ $\text{CuO} + \underline{2}$ $\text{NH}_3 \rightarrow \underline{3}$ $\text{Cu} + \underline{3}$ $\text{H}_2\text{O} + \underline{\quad}$ N_2
- $15.$ $\underline{2}$ $\text{Pb}(\text{NO}_3)_2 \rightarrow \underline{2}$ $\text{PbO} + \underline{4}$ $\text{NO}_2 + \underline{\quad}$ O_2
- $16.$ $\underline{\quad}$ $\text{Al}_2\text{O}_3 + \underline{3}$ $\text{C} + \underline{3}$ $\text{Cl}_2 \rightarrow \underline{3}$ $\text{CO} + \underline{2}$ AlCl_3
- $17.$ $\underline{2}$ $\text{Al} + \underline{2}$ $\text{H}_3\text{PO}_4 \rightarrow \underline{3}$ $\text{H}_2 + \underline{2}$ AlPO_4
- $18.$ $\underline{\quad}$ $\text{Al}_2(\text{SO}_4)_3 + \underline{6}$ $\text{NH}_4\text{Br} \rightarrow \underline{2}$ $\text{AlBr}_3 + \underline{3}$ $(\text{NH}_4)_2\text{SO}_4$
- $19.$ $\underline{\quad}$ $\text{Ca}_3(\text{PO}_4)_2 + \underline{3}$ $\text{H}_2\text{SO}_4 \rightarrow \underline{3}$ $\text{CaSO}_4 + \underline{2}$ H_3PO_4
- $20.$ $\underline{3}$ $\text{BaBr}_2 + \underline{2}$ $\text{Na}_3\text{PO}_4 \rightarrow \underline{\quad}$ $\text{Ba}_3(\text{PO}_4)_2 + \underline{6}$ NaBr
- $21.$ $\underline{\quad}$ $\text{Ca}_3(\text{PO}_4)_2 + \underline{3}$ $\text{SiO}_2 + \underline{5}$ $\text{C} \rightarrow \underline{2}$ $\text{P} + \underline{3}$ $\text{CaSiO}_3 + \underline{5}$ CO
- $22.$ $\underline{3}$ $\text{Hg}(\text{OH})_2 + \underline{2}$ $\text{H}_3\text{PO}_4 \rightarrow \underline{\quad}$ $\text{Hg}_3(\text{PO}_4)_2 + \underline{6}$ H_2O
- $23.$ $\underline{\quad}$ $\text{Na}_2\text{CO}_3 + \underline{2}$ $\text{HCl} \rightarrow \underline{2}$ $\text{NaCl} + \underline{\quad}$ $\text{H}_2\text{O} + \underline{\quad}$ CO_2
- $24.$ $\underline{4}$ $\text{P} + \underline{5}$ $\text{O}_2 \rightarrow \underline{2}$ P_2O_5