

NAME: \_\_\_\_\_

**Reminders:**

1. In a neutral atom the number of protons equals the number of electrons.
2. An atom can NEVER gain or lose protons
3. The number of protons equals the atomic number

# Ion Practice Set

1. What is an ion?
2. What does the number next to the ions signify?

**Complete the following table, using the periodic table in the back of your book.**

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	NUMBER OF ELECTRONS LOST OR GAINED
ex	Fluorine	F <sup>-</sup>	9	10	gained one
1			53	54	
2			16		gained two
3	potassium				lost one
4		Ca <sup>+2</sup>			
5			35	36	
6		Sr <sup>+2</sup>			
7		H <sup>+</sup>			
8			8		gained two
9			12		lost two
10	aluminum			10	
11			34	36	
12		H <sup>-</sup>			
13	lithium				lost one
14		Rb <sup>+</sup>			
15			17	18	

NAME: \_\_\_\_\_

# Isotopes Practice Set

1. What is an isotope?
2. What does the number next to isotopes signify?
3. How can you tell isotopes apart?

**For each of the following isotopes, write the number of protons, neutrons, and electrons.**

	Chromium-58	Chromium-63
# of protons		
# of neutrons		
# of electrons		

	Carbon-12	Carbon-16
# of protons		
# of neutrons		
# of electrons		

	Nitrogen-15	Nitrogen-20
# of protons		
# of neutrons		
# of electrons		

	Sulfur-23	Sulfur-25
# of protons		
# of neutrons		
# of electrons		

**Fill in the isotope names and any missing information, including isotope numbers from the chart. Use your periodic table and the information provided.**

	Iodine-	Iodine-
# of protons		
# of neutrons	32	35
# of electrons		

	Iron-	Iron-
# of protons		
# of neutrons	27	30
# of electrons		

# of protons		
# of neutrons	113	111
# of electrons	55	

# of protons	32	
# of neutrons	30	32
# of electrons		