

Name

Fitzell

Period

Atomic Structure

You can become more familiar with the atomic structure of some common substances by completing the chart below. For each substance, you have been given enough information to fill in all the blanks.

Substance	Symbol	Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
Helium	He	2	4	2	2	2
Magnesium	Mg	12	24	12	12	12
Zinc	Zn	30	65	30	35	30
Bromine	Br	35	80	35	45	35
Aluminum	Al	13	27	13	14	13
Uranium	U	92	238	92	146	92
Sodium	Na	11	23	11	12	11
Krypton	Kr	36	84	36	48	36
Calcium	Ca	20	40	20	20	20
Lithium	Li	3	7	3	4	3
Tungsten	W	74	184	74	110	74
Xenon	Xe	54	133	54	79	54
Magnesium	Mg	12	24	12	12	12
Carbon	C	6	12	6	6	6
Nitrogen	N	7	14	7	7	7
Silver	Ag	47	108	47	61	47

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Ion Practice Set

- Reminders:**
1. In a neutral atom the number of protons equals the number of electrons.
 2. An atom can NEVER gain or lose protons
 3. The number of protons equals the atomic number

1. What is an ion? *A charged atom or molecule*
2. What does the number next to the ions signify? *The amount of positive or negative charge*

Complete the following table, using the periodic table in the back of your book.

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	NUMBER OF ELECTRONS LOST OR GAINED
ex	Fluorine	F ⁻	9	10	gained one
1	Iodine	I ⁻	53	54	gained 1
2	Sulfur	S ²⁻	16	18	gained two
3	potassium	K ⁺	19	18	lost one
4	Calcium	Ca ⁺²	20	18	lost 2
5	Bromine	Br ⁻	35	36	gained 1
6	Strontium	Sr ⁺²	38	36	lost 2
7	Hydrogen	H ⁺	1	0	lost 1
8	Oxygen	O ²⁻	8	10	gained two
9	Magnesium	Mg ²⁺	12	10	lost two
10	aluminum	Al ³⁺	13	10	lost 3
11	Selenium	Se ²⁻	34	36	gained 2
12	Hydrogen	H ⁻	1	2	gained 1
13	lithium	Li ⁺	3	2	lost one
14	Rubidium	Rb ⁺	37	36	lost 1
15	Chlorine	Cl ⁻	17	18	gained 1