

Name \_\_\_\_\_

Period \_\_\_\_\_

# Atomic Structure

You can become more familiar with the atomic structure of some common substances by completing the chart below. For each substance, you have been given enough information to fill in all the blanks.

Substance	Symbol	Atomic Number	Mass Number	Number of Protons	Number of Neutrons	Number of Electrons
Helium		2	4			
Magnesium		12			12	
	Zn	30	65			
	Br		80			35
Aluminum				13	14	
Uranium					146	92
Sodium		11			12	
	Kr				48	36
	Ca		40	20		
Lithium		3	7			
Tungsten			184		110	74
Xenon					79	54
	Mg		24			
Carbon		6			6	
Nitrogen		7	14			
Silver				47	61	

NAME: \_\_\_\_\_

# Ion Practice Set

**Reminders:**  
1. In a neutral atom the number of protons equals the number of electrons.  
2. An atom can NEVER gain or lose protons  
3. The number of protons equals the atomic number

1. What is an ion?
2. What does the number next to the ions signify?

*Complete the following table, using the periodic table in the back of your book.*

	ELEMENT NAME	ION SYMBOL	NUMBER OF PROTONS	NUMBER OF ELECTRONS	NUMBER OF ELECTRONS LOST OR GAINED
ex	Fluorine	F <sup>-</sup>	9	10	gained one
<b>_____</b>					
1			53	54	
2			16		gained two
3	potassium				lost one
4		Ca <sup>+2</sup>			
5			35	36	
6		Sr <sup>+2</sup>			
7		H <sup>+</sup>			
8			8		gained two
9			12		lost two
10	aluminum			10	
11			34	36	
12		H <sup>-</sup>			
13	lithium				lost one
14		Rb <sup>+</sup>			
15			17	18	